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(71)(72) Applicants and Inventors: HEATH, James, R. [US/US]; 405 Hilgard Avenue, Los Angeles, CA 90095 (US). LEFF, Daniel, V. [US/US]; 405 Hilgard Avenue, Los Angeles, CA 90095 (US).			
(74) Agent: KONDZELLA, Michael, A.; Poms, Smith, Lande & Rose, Suite 3800, 2029 Century Park East, Los Angeles, CA 90067 (US).			

(54) Title: ORGANICALLY-FUNCTIONALIZED MONODISPERSE NANOCRYSTALS OF METALS

(57) Abstract

Organically-functionalized metal nanoparticles are produced by mixing a metal precursor with an organic surface passivant and reacting the resulting mixture with a reducing agent to generate free metal while binding the passivant to the surface of the free metal to produce organically functionalized metal particles.

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ORGANICALLY-FUNCTIONALIZED MONODISPERSE NANOCRYSTALS OF METALS**FIELD OF THE INVENTION**

This invention relates to metal and metal alloy nanocrystals. In one of its more particular aspects, this invention relates to organically-functionalized monodisperse nanocrystals of metals and metal alloys and methods of preparation thereof.

BACKGROUND OF THE INVENTION

Preparations for nanoparticle metal and metal oxide hydrosols are known. The myriad methods for preparing these particles include, but are not limited to: (1) the synthesis of colloidal dispersions of various transition metals (Pt, Pd, Ir, Rh, Os, Au, Ag, Fe, Co, and the like) in aqueous media, stabilized by added polymers as protective colloids; (2) the synthesis of ultrasmall metal oxide particles by the combination of water and metal chlorides, hydroxides, or acetates, in aqueous media; (3) the synthesis of Ag nanoparticles by the reduction of Ag⁺ in aqueous media; (4) the formation of colloidal silver and gold in aqueous media by ultrasonic radiation; (5) the formation of colloidal gold in aqueous media by the reduction of a gold salt; and (6) the formation of colloidal platinum and palladium in aqueous media by synthetic routes analogous to those for preparing gold colloids. The fabrication of large metal cluster complexes with various stoichiometries and ligands (e.g. M₅L₁₂Cl_x, M = Rh, Ru, Pt, Au; L = PR₃, AsR₃; x = 6, 20; R = Ph, t-Bu) is also known.

Nanometer-scale crystallites of various metals and non-metals have received a great deal of attention in the past decade. For such crystallites, the electronic, thermodynamic, and chemical properties depend sensitively on size, shape, and surface composition; therefore, these materials have been marked for a number of technological applications ranging from chemical catalysis, photoelectronics, film growth seeding, electronic materials, reprography, xerography, electron microscopy, and others. A major challenge to this field in general and a great barrier to actualizing such applications of these novel particles is the complete control over particle size, morphology, and surface structure. The preparation and isolation of crystallites characterized by well-defined surface compositions, narrow size distributions, and uniform shape is paramount to their success as applied materials. In addition, most of these applications require that the particles be dispersible into some solvent, polymer, or other

matrix as a monodisperse (non-aggregated) colloid, and that they exhibit various chemical and thermal stabilities. Metal particles involved in these various technologies include ferromagnetic materials (e.g. Fe_2O_3 , Co), noble metals (Pd, Pt), coinage metals (Au, Ag), alloys of these metals (e.g. Co_xAu_y), oxides of these metals (e.g. Ag_2O), and others.

5 To date, there are only a few reports concerning the preparation of organically-functionalized metal nanoparticles. However, the products synthesized in these reports are marred by some or all of the following characteristics: 1) the resulting materials are x-ray amorphous (non-crystalline); 2) the resulting materials have poor surface compositions; 3) the resulting materials have poor solubility in aqueous and organic media; and 4) the resulting
10 materials have broad relative size distributions (mean diameter \pm 50 %). For example, organically-functionalized gold nanocrystals limited to certain "average" sizes, characterized by broad size distributions; and which are not soluble in aqueous media have been disclosed.

BRIEF STATEMENT OF THE INVENTION

The nanocrystal products of the present invention avoid the previously discussed
15 problems. The "as prepared" spherical nanoparticles are crystalline and are characterized by narrow relative size distributions (as low as \pm 10 %). More importantly, however, is the fact that they are functionalized with well-defined organic groups covalently bound to the particle surface. These organic functional groups may be altered both chemically and with respect to percent coverage on the surface. Such chemistry is tremendously useful in the following
20 ways: 1) the organic functional groups, coupled with the size-dependent curvature of a particle surface, yield varying solubilities as a function of particle size. This fact, along with the size-dependent interactions between particles allows the production of narrow size distributions (ideally to within \pm .3 nm) over the range of ~ 1.5 to 10 nm, using specially developed chemical techniques; 2) the particles may be made highly soluble in aqueous
25 media as well as a variety of organics (alkanes, aromatics, halogenated hydrocarbons, and the like). This leads to the ability to prepare particles on or inside various substrates and matrices (gels, polymers, glasses, porous materials, and the like), 3) the effect of fractional surface coverages on similar size particles can be used to sterically control chemical approach to a surface to give increased reaction selectivity in certain catalytic processes.

The method for the preparation and isolation of crystallites according to the present invention is characterized by well-defined surface compositions, narrow size distributions, and uniform shape. Also, two- and three-dimensional close-packed ordered arrays (superlattices) of these nanocrystals have been fabricated on half-micron length scales.

- 5 Catalytically active metals such as Pt, Pd, and Ag and non-catalytically active metals and alloys such as Au and Co_xAu_y , have been prepared.

The present invention provides techniques for the synthesis of various metallic nanocrystal materials in which the resultant particles are characterized by the following properties: (1) they are soluble and resolvable in various organic media, including organic
10 solutions containing dissolved polymers; (2) they are stable as powders or monodisperse (non-aggregated) colloids under ambient conditions for at least several days; (3) they are stable for months when stored under low temperature conditions as powders or monodisperse (non-aggregated) colloids in solution; (4) they can exist as monodisperse entities (when prepared as organic colloids) which can be readily separated into arbitrarily narrow size
15 distributions via various chemical and chromatographic techniques; (5) they can be prepared in at least gram quantities; (6) they may consist of a host of metallic elements prepared as either pure metal particles or alloys, synthesized from the combination of a host of specific metal-containing inorganic compounds, phase transfer catalysts, surface passivants, and reducing agents; (7) they are readily dispersed into various matrices or onto various substrates
20 (gels, polymers, glasses, alumina, silica, and the like); (8) can be arranged into two- and three-dimensional close-packed ordered arrays to form 'superlattices' exhibiting novel electronic properties dominated by single electron phenomena. Particles that have been prepared and that meet the above criteria include Au, Ag, Pt, Pd, and Co/Au (alloy). Particle sizes range from 1 - 20 nm diameter.

- 25 Besides the fact that the nanocrystals of the present invention have a host of materials properties not exhibited by any other particle system, a major advancement of the technology over existing prior art is the use of covalently-bound organic ligands, which form excellent kinetic and/or thermodynamically stable monolayers on the surfaces of the nanocrystals, as a route toward stabilizing these particles in solution. This enables chemically tailored
30 solubility, monodispersity, and size control to the final metal nanocrystal product. The

ligands in general have a chemical component which interacts with the metal nanocrystal surface, and a chemical component which interacts with the surrounding solvent, polymer, matrix, etc. Both components can be modified within the limits of chemical compatibility.

The invention is not limited to simple single component metallic systems. Indeed, for some systems it is difficult to chemically stabilize a bare metal particle through the use of covalent organic ligands. However, in that case, alloys of the metal may be made in which a second metal is included. Ideally, this second metal is characterized by a lower surface energy (so it coats the surface of the particle), and is itself readily stabilized by a certain organic ligand. For example, in the Co/Au system, Co has magnetic properties which lend it to potential technological applications. However, it is difficult to find ligands which chemically stabilize a Co nanocrystal. Such ligands are readily found for the Au system, and, furthermore, Au has a lower surface energy. Thus it is possible to prepare Co/Au alloyed nanocrystals which exhibit the ferromagnetic properties of Co but are characterized by the relatively simple surface chemistry of Au.

In general, the organically functionalized metal and metal alloy nanoparticles of the present invention are prepared by providing a solution or dispersion of a metal precursor, providing a solution of an organic surface passivant, mixing the metal precursor solution or dispersion with the organic surface passivant solution, reacting the resulting mixture with a reducing agent to reduce the metal precursor to free metal while concomitantly binding the organic surface passivant to the resulting free metal surface to produce organically functionalized metal or metal alloy nanoparticles having a particle diameter of 10 - 200 \AA .

In a preferred embodiment of the invention, an organic solution of a phase transfer agent is mixed with the metal precursor prior to mixing with the organic surface passivant.

DESCRIPTION OF PREFERRED EMBODIMENTS

An inorganic gold compound such as HAuCl₄ is dissolved in H₂O to generate a solution containing AuCl₄⁻ as the active metal reagent. AuCl₄⁻ is phase transferred from H₂O into an organic phase such as toluene using an excess of a phase transfer reagent or catalyst such as N(C₈H₁₇)₄Br. A stoichiometric amount of an alkylthiol such as C₆H₁₃SH dissolved in an organic solvent such as toluene is added to the organic phase. Excess reducing agent such as NaBH₄ is dissolved in H₂O, added to the organic mixture with rapid stirring, and allowed

to continue to stir for several hours. The aqueous layer is removed and discarded. The organic layer is passed through submicron filter paper (no material is removed, and all color passes through the paper). The organically-functionalized metal nanocrystals are precipitated using an alcohol solution such as ethanol kept at low temperature. The filtrate is washed with this
5 same alcohol. The particles are re-dissolved in an organic solvent such as toluene, re-precipitated, and re-washed. The particles are finally re-dissolved in an organic solvent such as hexane or toluene.

Au particles with one phase transfer reagent and an alkylamine as the surface passivant can be prepared using an alkylamine such as C₁₂H₂₅NH₂ or C₁₈H₃₅NH₂ as the
10 surface passivant rather than an alkylthiol.

Au particles with no phase transfer reagent and an alkylamine as the surface passivant can be prepared using an alkylamine such as C₁₂H₂₅NH₂ or C₁₈H₃₅NH₂ as the surface passivant rather than an alkylthiol, and no phase transfer reagent. A small amount of insoluble black solid particulate material is generated during the synthesis. This precipitate is removed
15 by filtration of the two-phase system with submicron filter paper. The precipitation of the organically-functionalized metal nanocrystals then proceeds in the same manner above.

Ag particles with one phase transfer reagent and an alkylthiol as the surface passivant can be prepared using an inorganic silver compound such as AgNO₃ or AgClO₄ H₂O as the metal source, which, when dissolved in H₂O, yields Ag⁺ as the active metal reagent.

Pt particles with one phase transfer reagent and an alkylamine as the surface passivant can be prepared using an alkylamine such as C₁₂H₂₅NH₂ or C₁₈H₃₅NH₂ as the surface passivant and an inorganic platinum compound such as H₂PtCl₆ 3H₂O as the metal source, which, when dissolved in H₂O, yields PtCl₆⁻² as the active metal reagent.

Pd particles with one phase transfer reagent and an alkylamine as the surface passivant can be prepared using an alkylamine such as C₁₂H₂₅NH₂ or C₁₈H₃₅NH₂ as the surface passivant and an inorganic palladium compound such as Na₂PdCl₆ 4H₂O as the metal source, which, when dissolved in H₂O, yields PdCl₆⁻² as the active metal reagent.

Co/Au alloy particles with two phase transfer reagents and an alkylthiol as the surface passivant can be prepared as follows.

An inorganic cobalt compound (here $\text{CoCl}_2 \cdot \text{H}_2\text{O}$) is dissolved in H_2O to generate a solution containing Co^{+2} as the active metal reagent. Co^{+2} is phase transferred from H_2O into an organic phase such as toluene using an excess of a phase transfer reagent or catalyst such as $(\text{C}_6\text{H}_5)_4\text{BNa}$. The aqueous layer is removed and the organic layer is washed with H_2O .

5 An inorganic gold compound such as HAuCl_4 is dissolved in H_2O to generate a solution containing AuCl_4^- as the active metal reagent. AuCl_4^- is phase transferred from H_2O into an organic phase such as toluene using an excess of a phase transfer reagent or catalyst such as $\text{N}(\text{C}_8\text{H}_{17})_4\text{Br}$. The aqueous layer is removed and the organic layer is washed with H_2O . The two organic solutions are combined to form a mixture of Co^{+2} and AuCl_4^- . A stoichiometric amount of an alkylthiol such as $\text{C}_{12}\text{H}_{25}\text{SH}$ dissolved in toluene is added to the organic mixture. Excess reducing agent such as NaBH_4 is dissolved in H_2O , added to the organic mixture with rapid stirring, and allowed to continue to stir for several hours. The aqueous layer is removed and discarded. The organic layer is passed through submicron filter paper (no material is removed, and all color passes through the filter paper). The organically-functionalized alloy nanocrystals are precipitated using an alcohol solution such as ethanol kept at low temperature. The filtrate is washed with this same alcohol. The particles are re-dissolved in an organic solvent such as toluene, re-precipitated, and re-washed. The particles are finally re-dissolved in an organic solvent such as hexane or toluene.

Solubilization of organically-functionalized nanocrystals in aqueous media can be 20 accomplished as follows.

The nanocrystals are first prepared according to one of the synthetic schemes described above. A concentrated solution (e.g., 6 mg/ml) of the particular nanocrystals is prepared in an organic solvent such as hexane to yield an intensely-colored (e.g., purple brown, etc.) solution. A separate solution consisting of a specific weight % of a soap or 25 detergent molecule in aqueous media is prepared. The term "soap" or "detergent" is general here and is taken to mean any molecule that has a polar (hydrophilic) ionic region and a nonpolar (hydrophobic) hydrocarbon region (e.g., a fatty acid, an alkali metal alkane sulfonate salt, etc.). When dissolved in aqueous media under the appropriate conditions, these soaps and detergents will form structures called micelles. A micelle is basically any 30 water-soluble aggregate, spontaneously and reversibly formed from amphiphile molecules.

These aggregates can adopt a variety of three-dimensional structures (e.g., spheres, disks, bilayers, etc.) in which the hydrophobic moieties are segregated from the solvent by self-aggregation. If the hydrophobic portion of the amphiphile is a hydrocarbon chain, the micelles will consist of a hydrocarbon core, with the polar groups at the surface serving to maintain solubility in water. A nonpolar substance is solubilized in the hydrophobic region of these micelle structures. This is precisely the mechanism by which the soap or detergent solution the organically-functionalized nanocrystals. A known amount of the nanocrystal solution is added to a known amount of the colorless soap solution, resulting in a two-layer mixture. This mixture is stirred vigorously for a period of at least 12 hours. The color of the organic solution is transferred to the soap solution, and this signifies the solubilization of the metal nanocrystals in the aqueous media. The result is an intensely-colored single-layer solution containing a small amount of bulk metal that precipitates during the solubilization process. This metal precipitate is removed by filtration with submicron filter paper. The entire above procedure can be repeated several times in order to repeatedly increase the concentration of the metal nanocrystals in the aqueous media.

One example of a potential application of these materials concerns metal-doped matrices such as metal-doped polymer films. Thin polymer films for example containing a high weight percent of metal particles may provide a route to materials with a unique combination of mechanical, dielectric, optical, electric, and even magnetic properties. However, narrow particle size distributions, coupled with uniform distribution of the particles throughout the polymer film is necessary to make these properties microscopically uniform throughout a macroscopic film. Fabricating such a film would involve preparing polymer/solvent/particle solutions with relatively high and adjustable particle/polymer weight ratios and with the particles existing as monodisperse entities in the solution. The polymer/particle thin film could then be prepared from the solution through various standard spin-coating or evaporation techniques. Other suitable matrices include sol-gels, alumina, and glassy carbon.

A second example of a potential application of these materials deals with using silver particles in reprography. There are a number of reprographic processes which have stages that intimately depend on the nucleation and growth of small silver particles. For example, small

silver particles form the amplification (latent image) center in conventional photographic processes. This latent image center, formed by the action of light on silver halide crystals, acquires catalytic properties that enable it to trigger the reduction of the entire silver halide crystal to metallic silver by the reducing agent of the developer. For this process, uniform 5 particle size distributions lead to uniform film quality, and small particle sizes lead to enhanced film resolution. Also, since gold is frequently used in small quantities as a sensitizer for photographic emulsions, the Au particles may be applicable here as well.

A third example of a potential application of these materials pertains to chemical 10 catalysis. In catalytic processes, the size and morphology of the particle is often of great concern as it determines catalytic reactivity and selectivity. Organically-functionalized nanometer-scale particles of catalytically-active metals have extremely high surface areas (a large number of catalytically active sites per particle) and unique size-dependent chemical behavior which enables their application as highly selective catalysts in a variety of homogeneous and heterogeneous catalytic processes from petroleum cracking to polymer 15 synthesis. The Pt, Pd, and Ag particles are applicable here.

Other examples of potential applications include the use of these nanocrystals as 20 functional units in innovative micro and nanoelectronic devices. These applications are based on the idea that two- and three-dimensional close-packed ordered arrays (superlattices) of these nanocrystals will exhibit novel electronic properties dominated by single electron phenomena, due to the quantum confined electronic properties of the individual particles as well as their collective coherence effects.

The following examples illustrate specific embodiments of the present invention. In 25 the following examples all reactions were performed at room temperature, ambient pressure, and ambient atmosphere.

25

EXAMPLE 1

(a) 150 mg (.380 mmol) of $\text{HAuCl}_4 \cdot 3\text{H}_2\text{O}$ was dissolved by stirring in 25 mL of deionized water to yield a clear, yellow solution;

(b) 0.365g (.667 mmol) of $\text{N}(\text{C}_8\text{H}_{17})_4 \text{Br}_{(3)}$ was dissolved by stirring in 25 mL of toluene to yield a clear solution and then added to the rapidly-stirring aqueous solution of the 30 Au salt (solution (a)). An immediate two-layer separation resulted, with an orange/red organic

phase on top and an orange-tinted aqueous phase on the bottom. This mixture is vigorously stirred until all color disappeared from the aqueous phase, indicating quantitative transfer of the AuCl₄ moiety into the organic phase;

(c) 0.0190 g (.0226 ml; .108 mmol) of C₁₀H₂₁SH was placed in 25 mL of toluene and
5 then this mixture was added to the rapidly stirring two-phase mixture from (a) and (b);

(d) 0.151 g (4.00 mmol) of NaBH₄ was dissolved in 25 mL of deionized water to yield an effervescent, cloudy solution and then this mixture was added to the rapidly stirring mixture from (a), (b), and (c). There was an instant color change of the organic phase to black/brown and then quickly (1 minute) to dark purple. After 10 minutes, the aqueous layer
10 became clear and colorless. The reaction was continued at room temperature and room pressure (kept open to ambient atmosphere) for \approx 12 hour while rapidly stirring. Once the reaction time was finished, the aqueous phase was separated and discarded, and the dark purple organic phase was reduced in volume to \approx 5 mL by rotary evaporation. To this 5 ml toluene/particle solution was added 350 mL of methanol and this mixture was cooled to
15 -60°C for twelve hours. The dark purple/black precipitate was then vacuum filtered using 0.65 μ m nylon filter paper, washed with an excess of methanol (200 ml), and dried on a vacuum line to give \approx 60 mg of dry product. This 60 mg of particles was re-dissolved in 50 ml of toluene, re-precipitated, and re-washed by the procedure described just previously, to yield 40 mg of dry product. The particles were finally either stored as a powder in the freezer
20 or at room temperature, or they were re-dissolved in a preferred amount of an organic solvent (e.g., hexane, toluene, chloroform, etc.) to yield a solution with a concentration ranging from 1-30 mg/ml. These solutions were either stored in the freezer or at room temperature.

The nanoparticles were characterized by the following: (a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were
25 crystalline with diffraction peaks like those of fcc Au (except for the broadening at finite size). The main reflections were: (111) at 2 Θ = approx. 64.6°, (311) at 2 Θ = approx. 38.2°, (200) at 2 Θ = approx. 44.4°, (220) at 2 Θ = approx. 64.6°, (311) at 2 Θ = approx. 77.5°, (222) at 2 Θ = approx. 81.8°. Also, using diffraction peak line-width broadening, the average domain size was determined to be 70 \pm 7 Å; (b) Ultraviolet-visible spectroscopy (UV/vis):
30 This characterization, performed on dilute hexane or toluene solutions of the nanoparticles,

showed one main, broad absorption feature at $\lambda_{\max} = 521$ nm; (c) infrared spectroscopy (IR): This characterization, performed on a film of solid particles that were deposited on an NaCl window by evaporation of several drops of a particle/hexane solution, showed the standard C-C and C-H stretches, as well as those for the thiol group. The stretches were in the regions of 5 2950-2750 cm^{-1} , and 750-650 cm^{-1} ; (d) Nuclear magnetic resonance spectroscopy (NMR): This characterization, performed on concentrated particle/CDCl₃ solutions (10 mg/ml), showed three broad multiplets at $\delta = 1.50, 1.30$, and .90 ppm, with intensities of roughly 2.2:1. These peaks are superimposed on a fourth, very broad signal in the range of $\delta = 2.1 - .60$ ppm; (f) Transmission electron microscopy (TEM): This characterization, performed on 10 samples prepared by evaporating a drop of a dilute particle/hexane solution onto an amorphous carbon-coated Cu TEM grid, yielded TEM micrographs of the particles which indicated that the particles were predominantly spherical in morphology, that they were present with a broad size distribution ($\sigma \approx 20\%$), and that the average domain size was ≈ 65 Å; (g) X-ray photoelectron spectroscopy (XPS): This characterization, performed on a 15 uniform film of nanoparticles (several micrometers thick) supported on nylong filter paper, showed the appropriate signals for gold (5p_{3/2}, 4f_{7/2}, 4f_{5/2}, 4d_{5/2}, 4d_{3/2}, and 4p_{3/2} at $\approx 59, 84, 87, 336, 355$, and 548 eV, respectively), carbon (1s at ≈ 285.3 eV), and Oxygen (1s at ≈ 531.8 eV). Also observed were signals for Br (3p_{3/2} peak at 183.5 eV, 3p_{1/2} peak at 189.5 eV, and 3d peak at ≈ 68.0 ev). The peak positions, line shapes, and peak-to-peak distance of the Au 20 4f doublet are the standard measure of the gold oxidation state. The ginding energies for the Au 4f doublet are 83.5(3) and 87.2(3) eV (peak-to-peak distance of 3.7 eV). These measurements are consistent with the Au⁰ oxidation state; (h) Elemental analysis (EA): The analyses yielded 77.06 % Au, 2.99 % S, 2.86 % H, and 17.14 % C. The corresponding Au: S 25 molar ratio of the nanoparticles was 4.20:1, and the C:H and C:S ratios are those of neat decanethiol, within experimental uncertainties; (i) Differential scanning calorimetry (DSC): This characterization, performed on a 6 mg sample (dry powder) of nanoparticles, showed a broad, endothermic transition beginning at $\approx 95^\circ \text{C}$ and peaking at 120 ° C (18 J/g.); (j) Thermogravimetric analysis (TGA): This characterization, performed on a 5 mg sample (dry powder) of nanoparticles, showed a maximal rate of weight loss at approximately 235° C. 30 The total weight loss was found to be consistent with the total amount of bonded ligands

found by elemental analysis; (k) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility in hexane, toluene, chloroform, dichloromethane, pyridine, benzene, and several other organic solvents. Maximum solubility was found to be in the range 20 - 30 mg/ml.

5

EXAMPLE 2

- (a) 112 mg (.284 mmol) of HAuCl₄·3H₂O₍₃₎ was dissolved by stirring in 25 mL of deionized water to yield a clear, yellow solution;
- (b) 0.363 g (.666 mmol) of N(C₈H₁₇)₄Br₍₂₎ was dissolved by stirring in 25 mL of toluene to yield a clear solution and then added to the rapidly-stirring aqueous solution of the Au salt (solution (a)). An immediate two-layer separation resulted, with an orange/red organic phase on top and an orange-tinted aqueous phase on the bottom. This mixture is vigorously stirred until all color disappeared from the aqueous phase, indicating quantitative transfer of the AuCl₄ moiety into the organic phase;
- (c) 0.574 g (3.10 mmol) of C₁₂H₂₅NH₂₍₅₎. (dodecylamine) was placed in 25 mL of toluene and then this mixture was added to the rapidly stirring two-phase mixture from (a) & (b). Upon the addition of this solution, the aqueous layer immediately became beige/murky white;
- (d) 0.165g (4.86 mmol) of NaBH₄ was dissolved in 25 mL of deionzaed water to yield an effervescent, cloudy solution and then this mixture was added to the rapidly stirring mixture from (a), (b), and (c). There was an instant color change of the organic phase to black/brown and then quickly (1 minute) to dark purple. After 10 minutes, the aqueous layer became clear and colorless. The reaction was continued at room temperature and room pressure (kept open to ambient atmosphere) for ≈ 12 hour while rapidly stirring. Once the reaction time was finished, the aqueous phase was separated and discarded, and the dark purple organic phase was reduced in volume to ≈ 5 mL by rotary evaporation. To this 5 mi toluene/particle solution was added 350 mL of methanol and this mixture was cooled to -60° C for twelve hours. The dark purple/black precipitate was then vacuum filtered using 0.65 μm nylon filter paper, washed with an excess of methanol (200 ml), and dried on a vacuum line to give ≈ 60 mg of dry product. This 60 mg of particles was re-dissolved in 50 ml of toluene, re-precipitated, and re-washed by the procedure described just previously, to yield 60

mg of dry product. The particles were finally either stored as a powder in the freezer or at room temperature, or they were re-dissolved in a preferred amount of an organic solvent (e.g., hexane, toluene, chloroform, etc.) to yield a solution with a concentration ranging from 1-30 mg/ml. These solutions were either stored in the freezer or at room temperature. When stored 5 as powders at room temperature, the particles exhibit a certain degree a metastability. That is, the particles are unstable with respect to particle aggregation and quickly lose their solubility over a matter of a few days.

The nanoparticles were characterized by the following: (a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were 10 crystalline with diffraction peaks like those of fcc Au (except for the broadening at finite size). The main reflections were: (111) at 2Θ approx. 38.2° , (200) at Θ = approx. 44.4° , (220) at 2Θ approx. 64.6° , (311) at 2Θ = approx. 77.5° , (222) at 2Θ = approx. 81.8° . Also, using diffraction peak line-width broadening, the average domain size was determined to be $26 \pm 3 \text{ \AA}$; (b) Ultraviolet-visible spectroscopy (UV/vis): This characterization, performed on 15 dilute hexane or toluene solutions of the nanoparticles, showed one main, broad absorption feature at $\lambda_{\max} = 517 \text{ nm}$; (c) infrared spectroscopy (IR): This characterization, performed on a film of solid particles that were deposited on an NaCl window by evaporation of several drops of a particle/hexane solution, showed dodecylamine bands in the regions from $\approx 3310 - 2990 \text{ cm}^{-1}$ (N-H stretch), $\approx 3000-2850 \text{ cm}^{-1}$ (C-H allphatic stretch), $\approx 1700-1300 \text{ cm}^{-1}$ (N-H 20 band @ 1600 cm^{-1} and CH_2 scissor @ 1450 cm^{-1}), $\approx 1100-1050 \text{ cm}^{-1}$ (C-N stretch), and $\approx 900-700 \text{ cm}^{-1}$ (N-H wag); (d) Nuclear magnetic resonance spectroscopy (NMR): This 25 characterization, performed on concentrated particles/CDCl₃ solutions (10 mg/ml), showed three broad multiplets at $\delta = 1.56, 1.35$, and 85 ppm , with intensities of roughly 2:2:1. These peaks are superimposed on a fourth, very broad signal in the range of $\delta = 2.0 - .50 \text{ pp}$; (e) Mass spectroscopy (MS): This characterization, performed on solid samples, showed the typical fragmentation pattern of straight-chain primary amines as well as molecular ion peaks 30 of the amines. MS ($\text{Au}_x\text{dodecylamine}_y$),m/e (%): 30(100%) $[-\text{Cl}_2 \text{ NH}_2]^+$, 185 (M^+ , 4%) $[\text{C}_{12}\text{H}_{27}\text{N}]^+$; (f) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute particle/hexane solution onto an amorphous carbon-coated Cu TEM grid, yielded TEM micrographs of the particles which

indicated that the particles were predominantly spherical in morphology, that they were present with a broad size distribution ($\sigma \approx 20\%$), and that the average domain size was ≈ 30 Å; (g) X-ray photoelectron spectroscopy (XPS): This characterization, performed on a uniform film of nanoparticles (several micrometers thick) supported on nylon filter paper, showed the appropriate signals for gold ($5p_{3/2}$, $4f_{7/2}$, $4f_{5/2}$, $4d_{5/2}$, $4d_{3/2}$, and $4p_{3/2}$ at ≈ 59 , 84 , 87 , 336 , 366 , and 548 eV, respectively), carbon (1s at ≈ 285.3 eV), and Oxygen (1s at ≈ 531.8 eV). Also observed were signals for Br ($3p_{3/2}$ peak at 183.5 eV, $3p_{1/2}$ peak at 189.5 eV, and 3d peak at ≈ 68.0 ev). The peak positions, line shapes, and peak-to-peak distance of the Au 4f doublet are the standard measure of the gold oxidation state. The binding energies for the Au 4f doublet are $83.5(3)$ and $87.2(3)$ eV (peak-to-peak distance of 3.7 eV). These measurements are consistent with the Au° oxidation state; (h) Elemental analysis (EA): The analyses yielded 89.12% Au, $.79\%$ N, 2.00% H, and 9.20% C. The corresponding Au:N molar ratio of the nanoparticles was $7.9:1$, and the C:H and C:N ratios are those of neat dodecylamine, within experimental uncertainties; (i) Differential scanning calorimetry (DSC): This characterization, performed on a 7 mg sample (dry powder) of nanoparticles, showed a broad, exothermic transition(s) extending from ≈ 50 °C to 130 °C, which includes a relatively sharp endothermic feature centered at 90 °C (7 J/g); (j) Thermogravimetric analysis (TGA): This characterization, performed on a 5 mg sample (dry powder) of nanoparticles, showed a maximal rate of weight loss at approximately 250 °C. The total weight loss was found to be consistent with the total amount of bonded ligands found by elemental analysis; (k) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility in hexane, toluene, chloroform, dichloromethane, pyridine, benzene, and several other organic solvents. Maximum solubility was found to be in the range of 22 - 30 mg/ml.

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EXAMPLE 3

The process of EXAMPLE 2 was repeated except that no phase transfer reagent was used, a small amount of insoluble black-solid particulate material was generated during the synthesis, and this precipitate was removed by filtration of the two-phase system with submicron filter paper just before the precipitation step. That is, the insoluble precipitate was removed by filtration of the two-phase system with .66 micron filter paper. The aqueous

phase was then separated and discarded, and the dark-purple organic phase was reduced in volume to \approx 5 mL by rotary evaporation. The particles were then precipitated, reprecipitated, and stored in the manner described in EXAMPLE 2.

5 Particle composition, size, and properties may be varied by means of the following changes: the variation of the metal precursor used, the variation of phase transfer reagents used or their omission from the synthetic procedure, the variation of one or more surface passivants used, the variation of the reducing agent used, or the variation of some of the reactant molar ratios, or any combination thereof.

The nanoparticles were characterized by the following:

- 10 10 (a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were crystalline with diffraction peaks like those of fcc Au (except for the broadening at finite size). The main reflections were: (111) at $2\Theta = \text{approx. } 38.2^\circ$, (200) at $2\Theta = \text{approx. } 44.4^\circ$, (220) at $2\Theta = \text{approx. } 64.6^\circ$, (311) at $2\Theta = \text{approx. } 77.5^\circ$, (222) at $2\Theta = \text{approx. } 81.8^\circ$. Also, using diffraction peak line-width/broadening, the average domain size was determined to be $55 \pm 7 \text{ \AA}$; (b) Ultraviolet-visible spectroscopy (UV/vis): This characterization, performed on dilute hexane or toluene solutions of the nanoparticles, showed one main, broad absorption feature at $\lambda_{\max} = 525 \text{ nm}$; (c) infrared spectroscopy (IR): This characterization, performed on a film of solid particles that were deposited on an NaCl window by evaporation of several drops of a particle/hexane solution, showed dodecylamine bands in the regions from $\approx 3310-2990 \text{ cm}^{-1}$ (N-H stretch), $\approx 3000-2850 \text{ cm}^{-1}$ (C-H aliphatic stretch), $\approx 1700-1300 \text{ cm}^{-1}$ (N-H bend @ 1600 cm^{-1} and CH_2 scissor @ 1450 cm^{-1}), $\approx 1100-1050 \text{ cm}^{-1}$ (C-N stretch), and $\approx 900-700 \text{ cm}^{-1}$ (N-H wag); (d) Nuclear magnetic resonance spectroscopy (NMR): This characterization, performed on concentrated particle/CDCl₃ solutions (10 mg/ml), showed three broad multiplets at δ 1.54, 1.32, and .85 ppm, with intensities of roughly 2:2:1. These peaks are superimposed on a fourth, very broad signal in the range of δ 2.0 - .50 ppm; (e) Mass spectroscopy (MS): This characterization, performed on solid samples, showed the typical fragmentation pattern of straight-chain primary amines as well as molecular ion peaks of the amines. MS (Au_xdodecylamine_y), m/e (%): 30(100%) [-CH₂NH₂]⁺, 185 (M⁺, 4%) [C₁₂H₂₇N]⁺; (f) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute

particle/hexane solution onto an amorphous carbon/coated Cu TEM grid, yielded TEM micrographs of the particles which indicated that the particles were predominantly spherical in morphology, that they were present with a broad size distribution ($\sigma \approx 20\%$), and that the average domain size was $\approx 50\text{ \AA}$; (g) X-ray photoelectron spectroscopy (XPS): This characterization, performed on a uniform film of nanoparticles (several micrometers thick) supported on nylon filter paper, showed the appropriate signals for gold ($5p_{3/2}$, $4f_{7/2}$, $4f_{5/2}$, $4d_{5/2}$, $4d_{3/2}$, and $4p_{3/2}$ at ≈ 59 , 84 , 87 , 336 , 366 , and 548 eV , respectively), carbon ($1s$ at $\approx 285.3\text{ eV}$), and Oxygen ($1s$ at $\approx 531.8\text{ eV}$). Signals for Br ($3p_{3/2}$ peak at 183.5 eV , $3p_{1/2}$ peak at 189.5 eV , and $3d$ peak at $\approx 68.0\text{ ev}$) were not observed. The peak positions, line shapes, and peak-to-peak distance of the Au $4f$ doublet are the standard measure of the gold oxidation state. The binding energies for the Au $4f$ doublet are $83.5(3)$ and $87.2(3)\text{ eV}$ (peak-to-peak distance of 3.7 eV). These measurements are consistent with the Au° oxidation state; (h) Elemental analysis (EA): The analyses yielded 90.58% Au, $.75\%$ N, 1.69% H, and 9.51% C. The corresponding Au:N molar ratio of the nanoparticles was $8.6:1$, and the C:H and C:N ratios are those of neat dodecylamine, within experimental uncertainties; (i) Differential scanning calorimetry (DSC): This characterization, performed on a 8 mg sample (dry powder) of nanoparticles, showed a strong, broad, exothermic transition beginning at $\approx 50\text{ }^{\circ}\text{C}$ with a relatively sharp, and relatively endothermic feature peaking near $110\text{ }^{\circ}\text{C}$ (4 J/g); (j) Thermogravimetric analysis (TGA): This characterization, performed on a 5 mg sample (dry powder) of nanoparticles, showed a maximal rate of weight loss at approximately $250\text{ }^{\circ}\text{C}$. The total weight loss was found to be consistent with the total amount of bonded ligands found by elemental analysis; (k) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility in hexane, toluene, chloroform, dichloromethane, pyridine, benzene, and several other organic solvents. Maximum solubility was found to be in the range of $22\text{-}30\text{ mg/ml}$.

EXAMPLE 4

(a) 547 mg of $(\text{C}_8\text{H}_{17})_4\text{NBr}$ (phase transfer reagent) was dissolved in 10 ml of toluene and sonicated for 2 minutes ;

(b) 119 mg of $\text{CoCl}_2\text{6H}_2\text{O}$ was dissolved in 15 ml of H_2O by sonication for 15 minutes ;

- (c) The toluene and aqueous solutions from steps (a) and (b), respectively, were combined and stirred together for 15 minutes, which resulted in a blue-colored toluene layer. The aqueous phase was then separated from the organic phase and discarded;
- (d) 98 mg of HAuCl₄ was dissolved in 15 ml H₂O and then mixed with a 137 mg (C₈H₁₇)₄NBr in 20 ml toluene solution. The AuCl₄ ions were transferred from the aqueous to the toluene phase (organic Phase color becomes red/orange) and then the aqueous phase was separated and discarded;
- (e) The two solutions of metal precursors (1:2 Au:Co molar ratio) in toluene (solutions from step (c) and (d)) were merged and stirred for 5 minutes;
- 10 (f) 0.36 ml of C₁₂H₂₅SH (surface passivant) was added to the toluene solution from (e) and stirred for 2 minutes. The mixture tured blue/gray in color;
- (g) A solution of 283 mg NaBH₄ (reducing agent) in 3 ml H₂O was added to the toluene phase from step (f) and the reaction was allowed to proceed for 6 hours while stirring. Then, the black-colored toluene phase was separated from the aqueous phase and rotary 15 evaporated down to 5 ml. The concentrated solution was put in a freezer for 12 hours and then filtered, while cold, to remove phase transfer reagent that had crystallized out of the organic phase solution. The nanoparticles, still dissolved in the organic phase, were then precipitated by the addition of 300 ml of methanol. The particles/toluene/methanol solution was sonnicated for 10 min and then filtered through 0.2 µm nylon filter paper. The filtrate 20 was clear and the particles were black. The weight of residue on the filter paper was 41 mg. This residue was re-dissolved in 5 ml toluene, and the solution was sonnicated for 15 minutes and filtered. Then, the particles were precipitated again (using 200 ml of methanol) and filtered. The weight of the re-soluble, final residue was 20 mg. The particles were finally either stored as a powder in the freezer or at room temperature, or they were re-dissolved in a 25 preferred amount of an organic solvent (e.g., hexane, toluene, chloroform, etc.) to yield a solution with a concentration ranging from 1-30 mg/ml. These solutions were either stored in the freezer or at room temperature.
- The nanoparticles were characterized by the following materials characterization techniques:

(a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were crystalline with diffraction peaks like those of fcc Au (except for the broadening at finite size). The main reflections were: (111) at $2\Theta \approx$ approx. 38.2° , (200) at $2\Theta \approx$ approx. 44.4° , (220) at $2\Theta \approx$ approx. 64.6° , (311) at $2\Theta \approx$ approx. 77.5° , (222) at $2\Theta \approx$ approx. 81.8° . Cobalt reflections were masked by those of gold. Also, using diffraction peak line-width broadening, the average domain size was determined to be 30 ± 5 Å; (b) Ultraviolet/visible spectroscopy (UV/vis); This characterization, performed on dilute hexane or toluene solutions of the nanoparticles, showed one main, broad absorption feature at $\lambda_{max} = 520$ nm; (c) Infrared spectroscopy (IR): This characterization, performed on a film of solid particles that were deposited on an NaCl window by evaporation of several drops of a particle/hexane solution, showed the standard C-C and C-H stretches, as well as those for the thiol group. The stretches were in the regions of 2950 - 2750 cm^{-1} , 1500 - 1200 cm^{-1} , and 750 - 650 cm^{-1} ; (d) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute particle/hexane solution onto an amorphous carbon-coated Cu TEM grid, yielded TEM micrographs of the particles which indicated that the particles were predominantly spherical in morphology, that they were present with a relatively narrow size distribution ($\sigma \approx 10\%$), and that the average domain size was ≈ 30 Å; (e) X-ray, photoelectron spectroscopy (XPS): This characterization, performed on a uniform film of nanoparticles (several micrometers thick) supported on nylon filter paper, showed the appropriate signals for gold ($5p_{3/2}$, $4f_{7/2}$, $4f_{5/2}$, $4d_{5/2}$, $4d_{3/2}$, and $4p_{3/2}$ at ≈ 59 , 84 , 87 , 336 , 366 , and 548 eV, respectively), carbon (1s at ≈ 285.3 eV), and Oxygen (1s at ≈ 531.8 eV). The peak positions, line shapes, and peak-to-peak distance of the Au 4f doublet are the standard measure of the gold oxidation state. The binding energies for the Au 4f doublet are $83.5(3)$ and $87.2(3)$ eV (peak-to-peak distance of 3.7 eV). These measurements are consistent with the Au oxidation state. Also observed were the signals for cobalt (3s at 57 ev; $2p_{3/2}$ and $2p_{1/2}$ at 779 ev and 794 ev, respectively) and sulfur ($2p_{3/2}$ and $2p_{1/2}$ at 163 ev and 164 ev, respectively). An analysis of the XPS data revealed that the Co/Au alloy was comprised of about 3% Co and 97% Au; (f) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility in hexane, toluene,

chloroform, dichloromethane, pyridine, benzene, and several other organic solvents. Maximum solubility was found to be in the range of 20 - 30 mg/ml.

EXAMPLE 5

- (a) 10g of DDAB was dissolved in 104 ml of toluene and sonnicated for 10 minutes;
- 5 (b) 119 mg of $\text{CoCl}_2\text{6H}_2\text{O}$ was dissolved in the DDAS/toluene solution and sonnicated for 5 hours to dissolve all of the Co salt in the toluene. The CoCl_2 /DDAB/toluene solution had a typical cobalt blue color;
- 10 (c) 98 mg HAuCl_4 was dissolved in 15 ml H_2O and mixed with a 137 mg $(\text{C}_8\text{H}_{17})_4\text{NBr}$ in 20 ml toluene solution. The AuCl_4^- ions were transferred from the aqueous to the toluene phase (organic phase color becomes red/orange) and then the aqueous phase was separated and discarded;
- 15 (d) The two solutions (from steps (b) and (c)) of metal precursors (1:2 Au:Co molar ratio) in toluene were merged and stirred for 5 minutes. The solution had a dark green color;
- (e) 0.18 ml of $\text{C}_{12}\text{H}_{25}\text{SH}$ (surface passivant) was added to the toluene solution from 15 (d) and stirred for 2 minutes. The solution turned blue again;
- (f) A solution of 283 mg NaBH_4 (reducing agent) in 3 ml H_2O was added to the toluene phase resulting from step (a), and the reaction was allowed to proceed for 5 hours while stirring. After 5 hours of reaction time, the toluene phase was diluted with 200 ml a toluene and washed with 500 ml of H_2O . A viscous, white DDAB/water emulsion was formed and allowed to precipitate out of the thiol-capped Au/Co particles/toluene solution. The black particle/toluene solution was then separated and rotary evaporated to a concentrated 10 ml solution. 500 ml of methanol was then added to precipitate the particles. The particles/toluene/methanol solution was sonnicated for 30 min and then filtered through a 0.2 μm nylon filter paper. The filtrate was clear and the particles were black. The weight of residue on the filter paper was 69 mg. The residue was re-dissolved in 100 ml of toluene by sonnication for 15 minutes and the solution was then filtered. 31 mg of the residue were not dissolved. The toluene solution was rotary evaporated down to 5 ml and the particles were precipitated again by addition of 300 ml of methanol and 15 minutes sonnication. After filtering, the weight of the resolvable, final residue was 21 mg. The particles were finally 25 either stored as a powder in the freezer or at room temperature, or they were re-dissolved in a 30

preferred amount of an organic solvent (e.g., hexane, toluene, chloroform, etc.) to yield solution with a concentration ranging from 1-30 mg/ml. These solutions were either stored in the freezer or at room temperature.

The nanoparticles synthesized by the above procedures were characterized by the
5 following materials characterization techniques:

- (a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were crystalline with diffraction peaks like those of fcc Au (except for the broadening at finite size). The main reflections were: (111) at $2\Theta = \text{approx. } 38.2^\circ$, (200) at $2\Theta = \text{approx. } 44.4^\circ$, (220) at $2\Theta = \text{approx. } 64.6^\circ$, (311) at $2\Theta = \text{approx. } 77.5^\circ$, (222)
10 at $2\Theta = \text{approx. } 81.8^\circ$. Cobalt reflections were masked by those of gold. Also, using diffraction peak line-width broadening, the average domain size was determined to be 15 ± 2 Å; (b) Ultraviolet-visible spectroscopy (UV/vis): This characterization, performed on dilute hexane or toluene solutions of the nanoparticles, showed one main, broad absorption feature at $\lambda_{\max} = 517 \text{ nm}$; (c) Infrared spectroscopy (IR): This characterization, performed on a film
15 of solid particles that were deposited on an NaCl window by evaporation of several drops of a particle/hexane solution, showed the standard C-C and C-H stretches, as well as those for the thiol group. The stretches were in the regions of $2950\text{-}2750 \text{ cm}^{-1}$, $1500\text{-}1200 \text{ cm}^{-1}$, and $750\text{-}450 \text{ cm}^{-1}$; (d) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute particle/hexane solution onto an
20 amorphous carbon/coated Cu TEM grid, yielded TEM micrographs of the particles which indicated that the particles were predominantly spherical in morphology, that they were present with a relatively narrow size distribution ($\sigma \approx 7\%$), and that the average domain size was $\approx 15 \text{ \AA}$; (e) X-ray photoelectron spectroscopy (XPS): This characterization, performed on a uniform film of nanoparticles (several micrometers thick) supported on nylon filter paper,
25 showed the appropriate signals for gold ($5p_{3/2}$, $4f_{7/2}$, $4f_{5/2}$, $4d_{5/2}$, $4d_{3/2}$, and $4p_{3/2}$ at ≈ 59 , 84 , 87 , 336 , 366 , and 548 eV , respectively), carbon ($1s$ at $\approx 285.3 \text{ eV}$), and Oxygen ($1s$ at $\approx 531.8 \text{ eV}$). The peak positions, line shapes, and peak-to-peak distance of the Au 4f doublet are the standard measure of the gold oxidation state. The binding energies for the Au 4f doublet are $83.5(3)$ and $87.2(3) \text{ eV}$ (peak-to-peak distance of 3.7 eV). These measurements are consistent
30 with the Au^0 oxidation state. Also observed were the signals for cobalt ($3s$ at 57 eV ; $2p_{3/2}$ and

2p_{1/2} at 779 ev and 794 ev, respectively) and sulfur (2p_{3/2} and 2p_{1/2} at 163 ev and 164 ev, respectively). An analysis of the XPS data revealed that the Co/Au alloy was comprised of about 2% Co and 98% Au; (f) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility in hexane, toluene, chloroform, dichloromethane, pyridine, benzene, and several other organic solvents. Maximum solubility was found to be in the range of 20 - 30 mg/ml.

EXAMPLE 6

(a) Dodecanethiol-functionalized Ag nanocrystals (average domain size of 3 nm) were first prepared according to the procedure of EXAMPLE 1, except that AgNO₃ was used as the metal source and dodecanethiol was used as the thiol;

(b) A 6 mg/ml solution of the Dodecanethiol-functionalized Ag nanocrystals was prepared by dissolving 24 mg of particles in 4 ml of hexane to yield an intensely-colored (dark brown) solution;

(c) A separate solution (micelle solution) consisting of 20 g of sodium dodecylsulfate (SDS) dissolved in 300 ml of deionized H₂O was prepared. This yielded a 6.25 weight percent solution of SDS in H₂O; (d) 1 ml of the 6 mg/ml Ag particle/hexane solution was added to 20 ml of the 6.25 weight percent solution of SDS in H₂O resulting in a two-layer mixture (organic layer on top and aqueous layer on the bottom). This mixture was stirred vigorously for a period of 6 hours. The dark-brown color of the organic solution is transferred to the aqueous micelle solution to yield an amber-colored single phase system (no two layer separation exists anymore). This signifies the solubilization of the metal nanocrystals in the aqueous media. As a by-product of this solubilization procedure, a small amount of bulk metal precipitates. This metal precipitate was removed by filtration with .65 micron nylon filter paper to yield 1 mg of black, insoluble particulate material. The entire above procedure was repeated several times in order to increase the concentration of the metal nanocrystals in the aqueous media. A concentration of .10 mg/ml (.01 wt. % Ag) was ultimately achieved here.

The aqueous solutions of nanoparticles were characterized by the following techniques: (a) Ultraviolet-visible spectroscopy (UV-vis): This characterization, performed on dilute particle/hexane/SDS/water solutions, showed one main, broad absorption feature at

$\lambda_{\text{max}} = 450$ nm (this represents the characteristic optical signature of monodisperse silver colloids); (b) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute particle/hexane/SDS/water solution onto an amorphous carbon-coated Cu TEM grid, yielded TEM micrographs of the particles which
5 indicated that the particles were present with the same structural properties (e.g., shape, size, and size distribution) as those of the original dodecanethiol/functionalized Ag nanocrystals used for solubilization. Specifically, this analysis showed that the particles were predominantly spherical in morphology, that they were present with a relatively narrow size distribution ($\sigma \approx 10\%$), and that the average domain size was ≈ 30 Å.

10 EXAMPLE 7

- (a) 225 mg (.510 mmol) of $\text{H}_2\text{PtCl}_6\text{H}_2\text{O}_{(6)}$ was dissolved by stirring in 25 mL of deionized water to yield a clear, orange-yellow solution;
- (b) 0.620 g (1.13 mmol) of $\text{N}(\text{C}_8\text{H}_{17})_4\text{Br}_{(8)}$ was dissolved by stirring in 25 mL of toluene to yield a clear solution and then added to the rapidly-stirring aqueous solution of the
15 Pt salt (solution (a)). An immediate two-layer separation resulted, with an orange/red organic phase on top and an orange-yellow (tinted) aqueous phase on the bottom. This mixture is vigorously stirred until all color disappeared from the aqueous phase, indicating quantitative transfer of the PtCl_6^{2-} moiety into the organic phase;
- (c) .095 g (.511 mmol) of $\text{C}_{12}\text{H}_{25}\text{NH}_2_{(6)}$ (dodecylamine) was placed in 25 mL of
20 toluene and then this mixture was added to the rapidly stirring two-phase mixture from (a) and (b). Upon the addition of this solution, the aqueous layer immediately became beige/white; (d) 0.212 g (5.61 mmol) of NaBH_4 was dissolved in 25 mL of deionized water to yield an effervescent, cloudy solution and then this mixture was added to the rapidly stirring mixture from (a), (b) and (c). There was an instant color change of the organic phase to
25 black/brown and then quickly (1 minute) to dark brown. After 5 minutes, the aqueous layer became clear and colorless. The reaction was continued at room temperature and room pressure (kept open to ambient atmosphere) for ≈ 12 hour while rapidly stirring. Once the reaction time was finished, the aqueous phase was separated and discarded, and the dark-brown organic phase was reduced in volume to ≈ 5 mL by rotary evaporation. To this 5 ml
30 toluene/particle solution was added 350 mL of methanol and this mixture was cooled to -60°C

for twelve hours. The dark-brown precipitate was then vacuum filtered using 0.65 μm nylon filter paper, washed with an excess of methanol (220 ml), and dried on a vacuum line to give \approx 55 mg of dry product. This 55 mg of particles was re-dissolved in 50 ml of toluene, re-precipitated, and re-washed by the procedure described just previously, to yield 47 mg of dry product. The particles were finally either stored as a powder in the freezer or at room temperature, or they were re-dissolved in a preferred amount of an organic solvent (e.g., hexane, toluene, chloroform, and the like) to yield a solution with a concentration ranging from 1-30 mg/ml. These solutions were either stored in the freezer or at room temperature.

The nonoparticles characterized by the following:

- (a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were crystalline with diffraction peaks like those of fcc Pt (except for the broadening at finite size). The main reflections were: (111) at $2\Theta = \text{approx. } 38.2^\circ$, (200) at $2\Theta = \text{approx. } 44.4^\circ$, (220) at $2\Theta = \text{approx. } 64.6^\circ$, (311) at $2\Theta = \text{approx. } 77.5^\circ$, (222) at $2\Theta = \text{approx. } 81.8^\circ$. Also, using diffraction peak line-width broadening, the average domain size was determined to be $30 \pm 4\text{\AA}$;
- (b) Ultraviolet-visible spectroscopy (UV-vis): This characterization, performed on dilute hexane or toluene solutions of the nanoparticles, did not show an absorption feature in the visible spectrum between 300 - 800 nm (this is as expected because Pt is not a 'one-electron' metal);
- (c) infrared spectroscopy (IR): This characterization, performed on a film of solid particles that were deposited on an NaCl window by evaporation of several drops of a particle/hexane solution, showed dodecylamine bands in the regions from $\approx 3310\text{-}2990\text{ cm}^{-1}$ (N-H stretch), $\approx 3000\text{-}2850\text{ cm}^{-1}$ (C-H aliphatic stretch), $\approx 1700\text{-}1300\text{ cm}^{-1}$ (N-H bend @ 1600 cm^{-1} and CH_2 scissor @ 1450 cm^{-1}), $\approx 1100\text{-}1050\text{ cm}^{-1}$ (C-N stretch), and $\approx 900\text{-}700\text{ cm}^{-1}$ (N-H wag);
- (d) Nuclear magnetic resonance (NMR): This characterization, performed on concentrated particle/CDCl₃ solutions (10 mg/ml), showed three broad multiplets at $\delta = 1.56$, 1.34, and .87 ppm, with intensities of roughly 2:2:1. These peaks are superimposed on a fourth, very broad signal in the range of $\delta = 2.1$ - .55 ppm;
- (e) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute particle/hexane solution onto an amorphous carbon-coated Cu TEM grid, yielded TEM micrographs of the particles which indicated that the particles were predominantly spherical

in morphology, that they were present with a relatively narrow size distribution ($\sigma \approx 15\%$), and that the average domain size was $\approx 26\text{\AA}$; (f) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility hexane, toluene, chloroform, dichloromethane, pyridine, benzene, and several other organic solvents.

- 5 Maximum solubility was found to be in the range of 25-30 mg/ml.

EXAMPLE 8

- (a) 197 mg (.450 mmol) of $\text{Na}_2\text{PdCl}_8\text{H}_2\text{O}_{(8)}$ was dissolved by stirring in 25 mL of deionized water to yield a clear, gray/black solution;
- (b) 0.494 g (.900 mmol) of $\text{N}(\text{C}_8\text{H}_{17})_4\text{Br}_{(8)}$ was dissolved by stirring in 25 mL of toluene to yield a clear solution and then added to the rapidly-stirring aqueous solution of the Pd salt (solution (a)). An immediate two-layer separation resulted. This mixture is vigorously stirred until all color disappeared from the aqueous phase, indicating quantitative transfer of the PtCl_8^{-2} moiety into the organic phase (black);
- (c) .086 g (.465 mmol) of $\text{C}_{12}\text{H}_{25}\text{NH}_2_{(8)}$ (dodecylamine) was placed in 25 mL of toluene and then this mixture was added to the rapidly stirring two-phase mixture from (a) & (b). Upon the addition of this solution, the aqueous layer immediately became beige/white;
- (d) 0.171 g (4.52 mmol) of NaBH_4 was dissolved in 25 mL of deionized water to yield an effervescent, cloudy solution and then this mixture was added to the rapidly stirring mixture from (a), (b), and (c). There was an instant color change of the organic phase to dark black. After 5 minutes, the aqueous layer became clear and colorless. The reaction was continued at room temperature and room pressure (kept open to ambient atmosphere) for ≈ 12 hour while rapidly stirring. Once the reaction time was finished, the aqueous phase was separated and discarded, and the dark-black organic phase was reduced in volume to ≈ 5 mL by rotary evaporation. To this 5 ml toluene/particle solution was added 350 mL of methanol and this mixture was cooled to -60°C for twelve hours. The dark-black precipitate was then vacuum filtered using 0.65 μm nylon filter paper, washed with an excess of methanol (200 ml), and dried on a vacuum line to give ≈ 50 mg of dry product. This 50 mg of particles was re-dissolved in 50 ml of toluene, re-precipitated, and re-washed by the procedure described just previously, to yield 39 mg of dry product. The particles were finally either stored as a powder in the freezer or at room temperature, or they were re-dissolved in a preferred amount

of an organic solvent (e.g., hexane, chloroform, etc.) to yield a solution with a concentration ranging from 1-30 mg/ml. These solutions were either stored in the freezer or at room temperature.

The nanoparticles were characterized by the following:

- 5 (a) X-ray diffraction (XRD): This characterization, performed on a powder of the particles, showed that the particles were crystalline with diffraction peaks like those of fcc Pd (except for the broadening at finite size). The main reflections were: (111) at $2\Theta = \text{approx. } 38.2^\circ$, (200) at $2\Theta = \text{approx. } 44.4^\circ$, (220) at $2\Theta = \text{approx. } 77.5^\circ$, (311) at $2\Theta = \text{approx. } 77.5^\circ$, (222) at $2\Theta = \text{approx. } 81.8^\circ$. Also, using diffraction peak line-width broadening, the average
10 domain size was determined to be $20 \pm 3 \text{ \AA}$; (b) Ultraviolet-visible spectroscopy (uV/vis): This characterization, performed on dilute hexane or toluene solutions of the nanoparticles, did not show an absorption feature in the visible spectrum between 300 - 800 nm (this is as expected because Pd is not a 'one-electron' metal); (c) Infrared spectroscopy (IR): This characterization, performed on a film of solid particles that were deposited on an NaCl
15 window by evaporation of several drops of a particle/hexane soution, showed dodecylamine bands in the regions from $\approx 3310 - 2990 \text{ cm}^{-1}$ (N-H stretch), $\approx 3000-2850 \text{ cm}^{-1}$ (C-H aliphatic stretch), $\approx 1700-1300 \text{ cm}^{-1}$ (N-H bend @ 1600 cm^{-1} CH₂ scissor @ 1450 cm^{-1}), $\approx 1100-1050 \text{ cm}^{-1}$ (C-H stretch), and $\approx 900-700 \text{ cm}^{-1}$ (N-H wag); (d) Nuclear magnetic resonance
20 spectroscopy (NMR): This characterization, performed on concentrated particle/CDCl₃ solutation (10 mg/ml), showed three broad multiplets at $\delta = 1.54$, 1.36, and 192 ppm, with intesities of roughly 2:2:1. These peaks are superimposed on a fourth, very broad signal in the range of $\delta = 2.1\text{--}6.0 \text{ ppm}$; (e) Transmission electron microscopy (TEM): This characterization, performed on samples prepared by evaporating a drop of a dilute particle/hexane solutation onto an amorphous carbon-coated Cu TEM grid, yielded TEM
25 micrographs of the particles which indicated that the particles were predominatly spherical in morphology, that they were present with a relatively narrow size distribution ($\sigma \approx 10\%$), and that the average domain size was $\approx 18 \text{ \AA}$; (f) Solubility tests: This characterization, performed on dry powder samples of nanoparticles yielded excellent solubility hexane, toluene, chloroform, dichloromethane, pyridine, benzene and several other organic solvents.
30 Maximum solubility was found to be in the range of 25 -3 0 mg/ml.

EXAMPLE 9

- (a) 10 mg of dodecanethiol-capped Ag nanocrystals prepared according to the procedure of Example 6 and having an average domain size of approximately 30 Å was added to 10 mg of polystyrene and then mixed with 2 ml of toluene (effectively a 50 % by weight 5 Ag film because the toluene evaporates during spin coating procedures);
- (b) the 2 ml mixture from (a) was then spin coated onto a glass substrate (which contained patterned Al electrodes) at a rate of 3600 RPM, to generate a thin metal nanocrystal-doped polymer thin film;
- (c) After evaporating a top Al electrode onto the thin film, dielectric, optical and film 10 thickness measurements were carried out.

The film thickness was measured by profilometry to be 20 µm. The dielectric measurements of the metal nanocrystal-doped polymer thin film yielded unique dielectric values as compared to the 'pure' or un-doped polymer. The dielectric characteristics for the un-doped polymer thin film were: a) dielectric constant = 2; b) breakdown voltage - 12 kv/mm. 15 The dielectric characteristics for the metal nanocrystal-doped polymer thin film were: a) dielectric constant - 15; b) breakdown voltage - 1.2 kv/mm. As can be seen, the dielectric constant of the doped film increases by about a factor of 10. There is one broad absorption feature at $\lambda_{max} = 465$ nm. This feature is shifted to the red of that expected for just the nanoparticles in solution themselves (i.e. particles not part of a doped polymer film).
20 The fabrication methods discussed above may be used to prepare a range of novel metal nanocrystal-doped polymer thin film structures utilizing a variety of polymers (e.g., polystyrene, PMMA, conducting polymers, etc.) and a single kind of metal nanocrystal, or any combination of different metal nanocrystals. These include, but are not limited to: organically-functionalized Ag nanocrystal single or multilayer films; organically- 25 functionalized Au nanocrystal single or multilayer films; organically-functionalized Pt nanocrystal single or multilayer films; organically-functionalized Pd nanocrystal single or multilayer films; organically-functionalized Au/Co nanocrystal single or multilayer films; any combination therein of the organically-functionalized metal nanocrystals (e.g., a multilayer structure with an Ag/Au/Ag nanocrystal configuration or Ag/Pt/Au nanocrystal configuration,

etc.); any variable stoichiometric combination of the organically-functionalized metal nanocrystals (e.g., a 20% Ag / 20% Au / 10% Pt nanocrystal / 50% polymer configuration).

EXAMPLE 10

(a) dodecylamine-capped Pt nanocrystals were prepared and characterized according to the procedure of Example 7. The particles used here for catalysis had an average domain size of approximately 25 Å; (b) 10 ml of a 10 mg/ml solution of dodecylamine-capped Pt nanocrystals prepared as in (a), above, was combined with 125 ml of clean hexane in a reaction flask. The air was purged from this mixture through three freeze-pump-thaw vacuum cycles over the course of 5 hours; (c) H_{2(g)} was bubbled through the evacuated solution in the reaction flask from (b) for 12 hours; (d) The H_{2(g)} flow was stopped, the initial pressure of H_{2(g)} in the reaction vessel was measured with an in-line mercury manometer (P_{H_{2(g)}} = 665 torr), and then the flask was completely sealed; (e) .5 ml of 1-hexene was injected into the reaction vessel via a 3 cc syringe that had been purged with H_{2(g)}; (f) The reaction mixture was vigorously stirred for 5 hours, and the pressure drop of H_{2(g)} was monitored over this time interval. Control experiments (*no Pt-catalyst was added*) were also carried out in an analogous fashion.

Where no Pt nanocrystals were added, the pressure of H_{2(g)} remains constant at the initial value of 662 torr over the course of 250 minutes. This indicates that no reaction took place (*i.e.*, no conversion of 1-hexene to hexane occurred). Where an aliquot of a hexane solution of dodecylamine-functionalized Pt nanocrystals (avg. diameter ≈ 25 Å) was initially added, the pressure of H_{2(g)} initially decreases in a logarithmic fashion and then reaches some steady-state value (the conversion of 1-hexene to hexane is presumably complete within about 75 minutes or the catalyst is used up).

Any 10-100Å Pt and Pd nanocrystals functionalized with amine surface groups that will bind to the particles such as dodecylamine, octadecylamine, or pyridine can be used to produce similar results.

Although the present invention has been described with reference to preferred embodiments, workers skilled in the art will recognize that changes may be made in form and detail without departing from the spirit and scope of the invention.

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What is claimed is:

1. A method for preparing organically functionalized nanoparticles of metals and metal alloys comprising the steps of:
 - (a) providing a solution or dispersion of a metal precursor;
 - (b) providing a solution of an organic surface passivant;
 - 5 (c) mixing said metal precursor with said organic surface passivant;
 - (d) reacting the resulting mixture with a reducing agent to reduce said metal precursor to free metal while concomitantly binding said organic surface passivant to the resulting free metal surface;
thereby providing organically functionalized metal and metal alloy nanoparticles having a
10 particle diameter of 10-200Å.
2. The method of claim 1 which additionally comprises the steps of providing an organic solution of a phase transfer agent and mixing said phase transfer agent with said metal precursor prior to mixing with said organic surface passivant.
3. The method of claim 2 wherein said metal precursor contains one or more metal
15 atoms.
4. The method of claim 3 wherein said one or more metal atoms is an alkali metal atom.
5. The method of claim 3 wherein said one or more metal atoms is an alkaline earth metal atom.
6. The method of claim 3 wherein said one or more metal atoms is a transition metal
20 atom.
7. The method of claim 6 wherein said transition metal atom is a member selected from the group consisting of Au, Ag, Pt, Pd, Co, and alloys thereof.

8. The method of claim 3 wherein said one or more metal atoms is a Group III metal atom.
9. The method of claim 8 wherein said one or more Group III metal atoms is a member selected from the group consisting of B, Al, Ga, In, and Tl.
- 5 10. The method of claim 3 wherein said one or more metal atoms is Pb.
11. The method of claim 2 wherein said phase transfer agent is a molecule containing both polar and non-polar functionality.
12. The method of claim 2 wherein said phase transfer agent is a molecule having the ability to form micelles or inverted micelles.
- 10 13. The method of claim 2 wherein said phase transfer agent is an amphiphilic molecule.
14. The method of claim 13 wherein said amphiphilic molecule is a member selected from the group consisting of alcohols, ethers, esters, fatty acids, phospholipids, polyphosphate esters, polyethers, alkylammonium salts, tetralkylboron alkali metal compounds, alkali metal soaps and detergents, and nitrogen-containing aromatic compounds.
- 15 15. The method of claim 2 wherein said phase transfer agent is a zwitterion molecule.
16. The method of claim 2 wherein said organic surface passivant is a compound of the formula R-X, wherein R is a member selected from the group consisting of alkyl, aryl, alkynyl and alkenyl groups, and X is a group which can bind to said free metal surface via strong or weak interactions.
- 20 17. The method of claim 2 wherein said organic surface passivant is a member selected from the group consisting of thiols, phosphines, oxyphosphines, disulfides, amines, oxides and amides.

18. The method of claim 2 wherein said reducing agent has the ability to reduce metal ions, metal complex ions and metal clusters to the metallic state.
19. The method of claim 2 wherein said reducing agent is a member selected from the group consisting of sodium borohydride, sodium cyanoborohydride, sodium citrate, lithium aluminum hydride, K and NaK.
 - 5
20. A method for preparing a metal nanoparticle-doped matrix or a metal alloy nanoparticle-doped matrix comprising:
 - providing organically functionalized metal or metal alloy particles prepared according to the method of claim 2;
 - 10 providing a matrix; and
 - combining said organically functionalized metal or metal alloy particles with said matrix.
21. The method of claim 20 wherein said matrix is a polymer solution.
22. The method of claim 21 wherein said polymer solution and said functionalized metal or metal alloy particles are combined by spin coating.
 - 15
23. The method of claim 21 wherein the polymer comprising said polymer solution is a member selected from the group consisting of polystyrene, polymethylmethacrylate, polyethers, polypropylene, and polyethylene.
24. The method of claim 21 wherein the solvent comprising said polymer solution is a member selected from the group consisting of alcohols, ketones, ethers, chloroform, TCE and dichloromethane.
 - 20
25. The method of claim 20 wherein said matrix is a sol-gel.
26. The method of claim 20 wherein said matrix is alumina.

27. The method of claim 20 wherein said matrix is glassy carbon.
28. The product of the method of claim 1.
29. The product of the method of claim 2.
30. A method for converting 1-hexene to hexane which comprises reacting 1-hexene
5 with hydrogen in the presence of a catalyst prepared by the steps of:
 - (a) providing a solution of a Pt or Pd salt;
 - (b) mixing said Pt or Pd salt solution with an organic solution of a phase transfer agent;
 - (c) providing a solution of an amine surface passivant selected from the group
10 consisting of dodecylamine, octadecylamine, and pyridine;
 - (d) mixing the organic phase resulting from step (b) with said amine surface passivant solution; and
 - (e) reacting the resulting mixture of step (d) with a reducing agent to reduce said Pt or Pd salt to free Pt or Pd while concomitantly binding said amine surface passivant to the
15 resulting free Pt or Pd surface.

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US96/20402

A. CLASSIFICATION OF SUBJECT MATTER

IPC(6) :B32B 1/00
 US CL : 428/402; 427/212

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 428/402; 427/212

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched
none

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

none

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X, P	US 5,585,457 A (NEWKOME ET AL.) 17 December 1996, column 15, lines 28-43; column 7, lines 14-48.	1-30
Y, E	US 5,620,584 A (REETZ ET AL.) 15 April 1997, column 2, lines 25-55.	1-30
X	US 4,152,303 A (COHEN et al.) 01 May 1979, column 2, lines 14-60.	1-30

<input type="checkbox"/>	Further documents are listed in the continuation of Box C.	<input type="checkbox"/>	See patent family annex.
*	Special categories of cited documents:	T	later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
"A"	document defining the general state of the art which is not considered to be of particular relevance	"X"	document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step
"E"	earlier document published on or after the international filing date	"Y"	document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is taken alone
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"P"	document published prior to the international filing date but later than the priority date claimed		

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